

UNIVERSITE DE NEUCHATEL
Faculté des Sciences
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**Clusters organométalliques amphiphiles :
synthèse, caractérisation et organisation supramoléculaire
par la technique de Langmuir-Blodgett**

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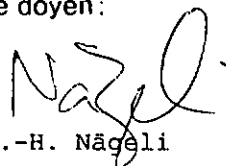
La Faculté des sciences de l'Université de Neuchâtel
sur le rapport des membres du jury,

MM. les professeurs R. Deschenaux, G. Süss-Fink,
N. de Rooij, R. Steiger et H. Kuhn (Göttingen)

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autorise l'impression de la présente thèse.

Neuchâtel, le 22 juin 1994.....

Le doyen :



H.-H. Nägeli

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An Amphiphilic Trinuclear Cobalt Cluster-containing Molecular Unit: Synthesis, Characterization and Langmuir–Blodgett Films†

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An amphiphilic polynuclear transition-metal cluster has been prepared and its capability for forming layers at the air–water interface and Langmuir–Blodgett films on different substrates investigated. The compound was obtained by linking a $\text{Co}_3(\text{CO})_9$ cluster to a hydrophobic cholesterol framework. A polar head was introduced by substituting one CO ligand by an isocyanide derivative bearing a hydrophilic functional group. The surface pressure and surface potential vs. area isotherms and Brewster-angle microscopy indicated the formation of stable Langmuir films. Up to thirty-six transfers were performed on hydrophobized solids (mica, glass, quartz and silicon wafers). Structural characterization of the transferred films (UV and IR spectroscopy, atomic force microscopy, small-angle X-ray scattering) demonstrated the formation of stable, highly organized crystals. The results represent a novel strategy towards the development of supramolecular assemblies from polynuclear transition-metal compounds and organometallic surfaces. The structure of $[\text{Co}_3(\text{CO})_8(\text{CCO}_2\text{Me})(\text{CNC}_6\text{H}_4\text{CN})]$ has been established by X-ray crystallography.

The Langmuir–Blodgett technique is currently the subject of intense investigation:¹ the possibility of fabricating tailor-made supramolecular assemblies, with well defined structures,² and of studying specific properties at the molecular level has allowed a better understanding of fundamental phenomena occurring at the frontiers of chemistry, biology and physics.³ From a practical point of view, this technique could lead to the development of new technologies, such as in the miniaturization of electronic and optical devices (nanometer-scale technology).

Recent studies have shown that a large variety of structures can be used as amphiphiles. Stable monolayers and Langmuir–Blodgett films have been obtained from porphyrins,⁴ phthalocyanins,⁵ oligothiophenes,⁶ polymers,⁷ ferrocenes,⁸ ruthenium complexes,⁹ cyclodextrins,¹⁰ macrocyclic polyamides¹¹ and calixarenes.¹² This clearly demonstrates the wide scope of the method. The design and the synthesis of new amphiphiles, combining an original and elegant structure with unique properties, constitutes a challenging goal towards the elaboration of new materials based on this technology.

In the present paper we describe the synthesis and characterization of an amphiphilic organometallic cluster derived from a tricobalt core, and demonstrate, for the first time, that polynuclear transition-metal frameworks can be used as molecular units to form stable Langmuir–Blodgett films. The remarkable magnetic,¹³ electrochemical¹⁴ and structural¹⁵ features of organometallic clusters motivated us to undertake this research.

Results and Discussion

Synthesis.—The investigated amphiphilic cluster **I** consists of a trinuclear cobalt core connected to a cholesterol (cholest-5-en-3 β -ol) moiety, acting as the hydrophobic tail. The nitrile

function, introduced by substitution of one CO ligand with an isocyanide derivative, acts as the polar headgroup. The synthesis of **I** is shown in Scheme 1. Treatment of methyl 4-hydroxybenzoate **I** with dimethoxymethane and P_2O_5 in CH_2Cl_2 gave the protected phenol intermediate **II** (89%), which was transformed into the cholesterol derivative **III** (84%) by transesterification [LiBu , cholesterol, tetrahydrofuran (thf)]. Removal of the protecting group under acidic conditions ($\text{HCl-MeCO}_2\text{H}$, thf) yielded **IV** (85%). Alkylation of the latter with 2-bromoethanol in the presence of K_2CO_3 gave **V** (83%). The nonacarbonyltricobalt cluster **2** was prepared, adapting a two-step literature procedure,¹⁶ in 44% yield from **3**¹⁷ and **V**: chloro cluster **3** was first treated with AlCl_3 in CH_2Cl_2 to produce an acylium intermediate,¹⁷ which was esterified with alcohol **V**. Reaction of **2** with isocyanide **VIII** (thf, room temperature) afforded, after purification by column chromatography (silica gel; hexane– CH_2Cl_2 – Et_2O , 6:2:1) and crystallization from hexane, the targeted cluster **1** (25%). The isocyanide **VIII** was synthesised as depicted in Scheme 2: 4-aminobenzonitrile **IX** was heated in HCO_2H to give **X** (75%), dehydration of which in $\text{POCl}_3\text{-NHPri}_2$ ¹⁸ led to **VIII** (80%) as a white solid.

Compound **VIII** is a potentially bidentate ligand. Its co-ordination to cluster **I** by the isocyanide function is demonstrated by comparing their IR spectra (KBr). Indeed, the wavenumber of the nitrile group is identical for both compounds (2232 cm^{-1} for **VIII** and 2233 cm^{-1} for **I**), while that of the isocyanide function is clearly different: 2133 cm^{-1} for unco-ordinated **VIII** and 2152 cm^{-1} for **I**. The observed shift of 19 cm^{-1} results unambiguously from the co-ordination of the isocyanide group.

Solid-state Structure.—To understand the behaviour and arrangement of the investigated amphiphile **I** at the air–water interface and within the transferred films its structure must be known. The monosubstitution was confirmed by ^1H NMR spectroscopy and elemental analysis. However, the substitution

† Supplementary data available: see Instructions for Authors, *J. Chem. Soc., Dalton Trans.*, 1994, Issue 1, pp. xxiii–xxviii.

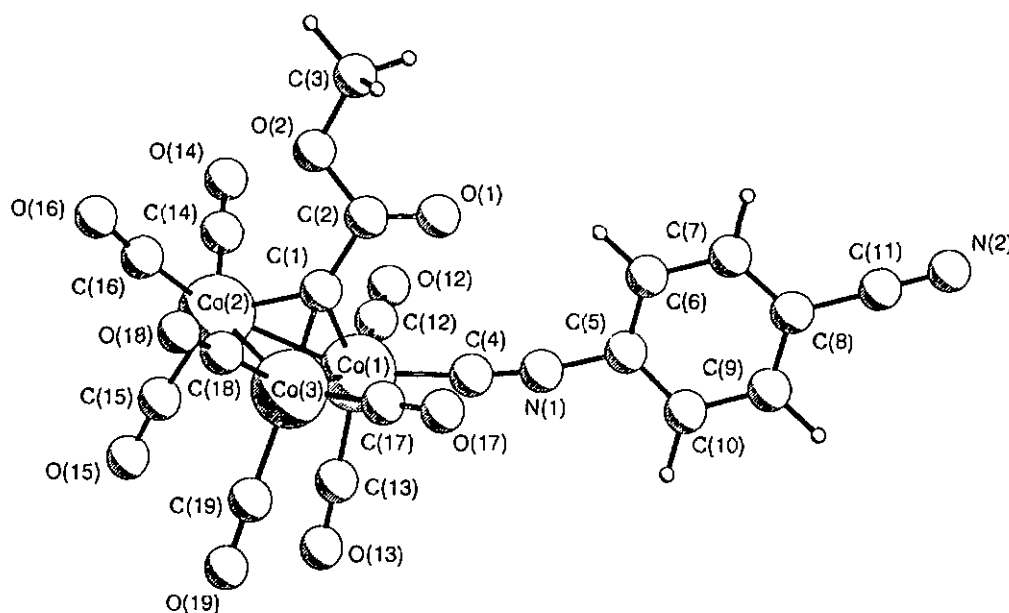
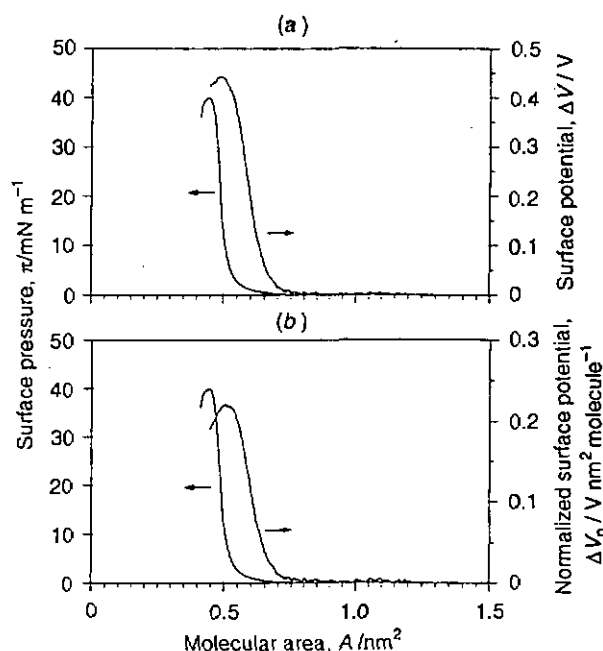
Fig. 1 A PLUTO¹⁹ plot of cluster 4 showing the arbitrary numbering scheme

Table 1 Atomic coordinates for cluster 4

Atom	x	y	z
Co(1)	0.013 36(5)	0.214 62(5)	0.174 27(3)
Co(2)	0.206 62(6)	0.022 14(6)	0.148 92(4)
Co(3)	0.285 84(5)	0.281 62(5)	0.117 54(3)
C(1)	0.171 4(4)	0.137 9(4)	0.240 5(3)
C(2)	0.174 4(5)	0.115 8(5)	0.348 9(3)
O(1)	0.127 3(5)	0.200 3(5)	0.396 6(3)
O(2)	0.241 9(5)	-0.015 6(4)	0.386 3(3)
C(3)	0.252 5(9)	-0.043 3(7)	0.492 4(4)
N(1)	-0.125 5(4)	0.423 7(4)	0.294 6(3)
N(2)	-0.513 1(6)	0.721 1(6)	0.662 4(4)
C(4)	-0.068 2(4)	0.350 4(5)	0.244 8(3)
C(5)	-0.200 2(4)	0.491 6(4)	0.369 4(3)
C(6)	-0.169 2(6)	0.441 9(5)	0.465 3(4)
C(7)	-0.247 0(5)	0.504 2(6)	0.540 5(4)
C(8)	-0.352 5(5)	0.616 7(5)	0.518 7(3)
C(9)	-0.379 5(5)	0.667 7(5)	0.421 7(3)
C(10)	-0.302 7(5)	0.606 1(5)	0.345 5(3)
C(11)	-0.440 6(5)	0.675 3(6)	0.598 9(4)
C(12)	-0.143 0(5)	0.086 9(5)	0.240 0(3)
O(12)	-0.241 5(4)	0.005 3(4)	0.283 4(3)
C(13)	-0.030 9(4)	0.295 2(4)	0.049 1(3)
O(13)	-0.056 6(4)	0.340 6(4)	-0.029 5(3)
O(14)	0.027 9(6)	-0.238 7(5)	0.294 4(5)
C(15)	0.193 3(6)	0.037 7(6)	0.019 8(4)
O(15)	0.177 9(5)	0.052 7(6)	-0.060 3(3)
C(16)	0.382 2(6)	-0.070 4(6)	0.161 8(5)
O(16)	0.488 3(5)	-0.136 3(5)	0.176 6(5)
C(17)	0.269 1(5)	0.438 8(5)	0.163 9(3)
O(17)	0.261 3(5)	0.537 3(4)	0.194 1(3)
C(18)	0.482 6(5)	0.239 9(5)	0.113 6(3)
O(18)	0.607 1(4)	0.215 0(5)	0.108 8(4)
C(19)	0.302 1(5)	0.370 3(5)	-0.021 9(3)
O(19)	0.311 4(5)	0.423 8(5)	-0.107 07(24)

Table 2 Selected bond lengths (nm) and angles (°) for cluster 4

Co(1)-Co(2)	0.246 43(8)	Co(1)-C(4)	0.185 8(4)
Co(1)-Co(3)	0.247 21(8)	Co(1)-C(12)	0.178 9(4)
Co(2)-Co(3)	0.248 30(9)	Co(1)-C(13)	0.182 2(4)
Co(1)-C(1)	0.190 2(4)	N(1)-C(4)	0.114 1(5)
Co(2)-C(1)	0.190 0(4)	N(1)-C(5)	0.139 9(5)
Co(3)-C(1)	0.190 8(4)	N(2)-C(11)	0.114 1(6)
Co(2)-Co(1)-Co(3)	60.397(24)	C(1)-Co(1)-C(12)	106.94(17)
Co(1)-Co(2)-Co(3)	59.956(24)	Co(2)-Co(1)-C(12)	95.78(13)
Co(1)-Co(3)-Co(2)	59.647(21)	Co(3)-Co(1)-C(4)	100.97(12)
C(1)-Co(1)-C(4)	98.17(17)		

Fig. 2 Surface pressure vs. area isotherms of compound V (20.0 °C, aqueous CdCl₂-NaHCO₃ subphase) with (a) surface potential and (b) normalized surface potential ($\Delta V_n = \text{surface potential} \times \text{molecular area}$)

showed the typical behaviour of amphiphiles at the air-water interface. Indeed, the surface pressure and surface potential vs. area isotherms (Fig. 2) clearly demonstrated the formation of a stable film at the air-water interface. The limiting area $A_0 = 0.52 \text{ nm}^2$ also confirmed the obtention of a crystalline solid. The results obtained for V are in agreement with literature data reported for other cholesterol derivatives.²¹

Organometallic-containing films were prepared by spreading the amphiphile 1 from CHCl₃ solutions on different subphases and at different temperatures. The π vs. A isotherms showed that

compound 1 formed stable films between 10 and 20 °C on a CdCl₂-NaHCO₃ solution (Fig. 3) as well as on pure water (Fig. 4). A strong reduction of the collapse pressure, corresponding

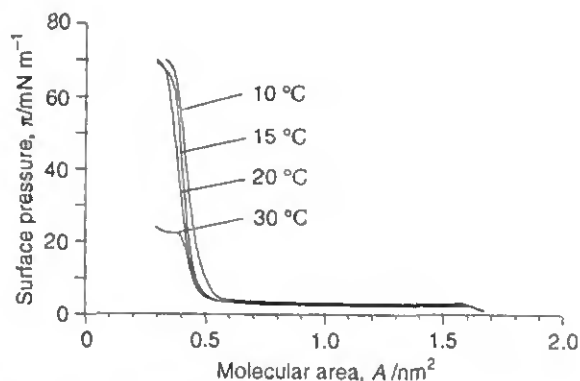


Fig. 3 Surface pressure vs. area isotherms of cluster 1 at different temperatures on an aqueous $\text{CdCl}_2\text{-NaHCO}_3$ subphase

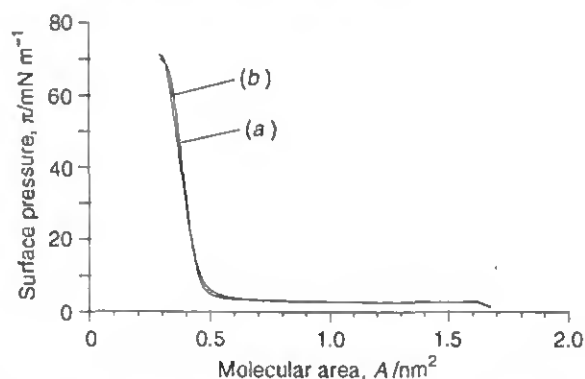


Fig. 4 Comparison of the surface pressure vs. area isotherms of cluster 1 on pure water (a) and on aqueous $\text{CdCl}_2\text{-NaHCO}_3$ subphases at 20.0°C

to a decrease in film stability, was observed at 30°C . At 20°C a limiting area $A_0 = 0.46 \text{ nm}^2$, which corresponds to approximately half the value calculated from the CSC Chem3D PlusTM models ($0.80\text{--}0.85 \text{ nm}^2$, see Fig. 10), was determined for each subphase. This finding indicated that amphiphile 1 did not organize in a classical monolayer at high surface pressure. Lowering of the temperature affected the limiting area only slightly: $A_0 = 0.47$ and 0.49 nm^2 at 15 and 10°C , respectively.

The surface potentials measured during the compression of cluster 1 (Fig. 5) revealed interesting features and suggested a head-to-tail organization of 1 [see Fig. 12(b)]. First, a sharp increase of ΔV , corresponding to the slight increase of the surface pressure, was observed at 1.80 nm^2 . This result was attributed to the formation of a monolayer of low stability (5 mN m^{-1}). Secondly, the regular decrease of the normalized surface potential [Fig. 5(b)], upon further compression, while the surface pressure remained constant (plateau), clearly indicated the transformation of the monolayer into a new structure in which the molecular units of 1 must be arranged in such a way as to produce an assembly with a low overall dipole moment. This situation is achieved when the molecular units of 1 are associated in pairs with the tricoalt cluster cores pointing in opposite directions, *i.e.* in a head-to-tail arrangement, which confirms the A_0 values obtained from the π vs. A isotherms.

Brewster-angle microscopy²² (BAM) was used to visualize the phenomena occurring at the air-water interface. Fig. 6 shows photomicrographs at different stages during the compression of cluster 1 on a $\text{CdCl}_2\text{-NaHCO}_3$ subphase at 20°C . Similar observations were obtained on pure water. Highly reflectant circular domains appeared at 1.60 nm^2 , the size and number of which increased upon compression in the plateau region [Fig. 6(a) and 6(b)]. From 0.50 nm^2 , *i.e.* during the increase of the surface pressure, the circular domains approached each other and assembled to produce a uniform surface [Fig. 6(c)]. The collapse of the latter is presented in Fig. 6(d). Several hysteresis curves were recorded between 0 and 60

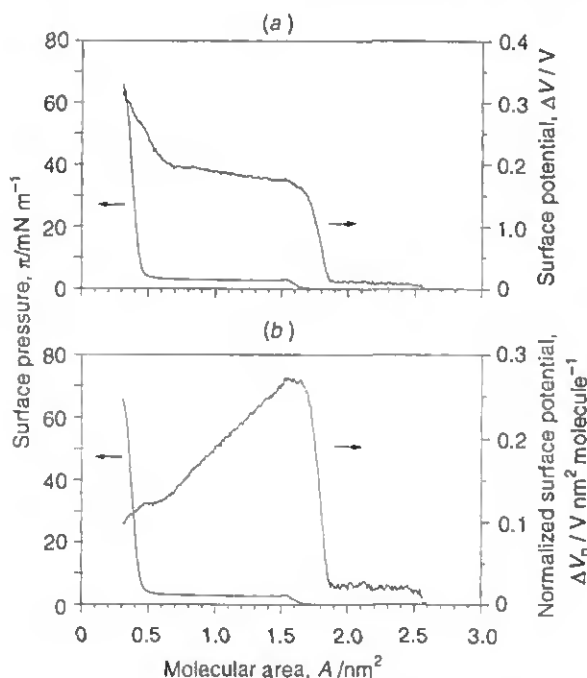


Fig. 5 Surface pressure vs. area isotherms of cluster 1 (20.0°C , aqueous $\text{CdCl}_2\text{-NaHCO}_3$ subphase) with (a) surface potential and (b) normalized surface potential

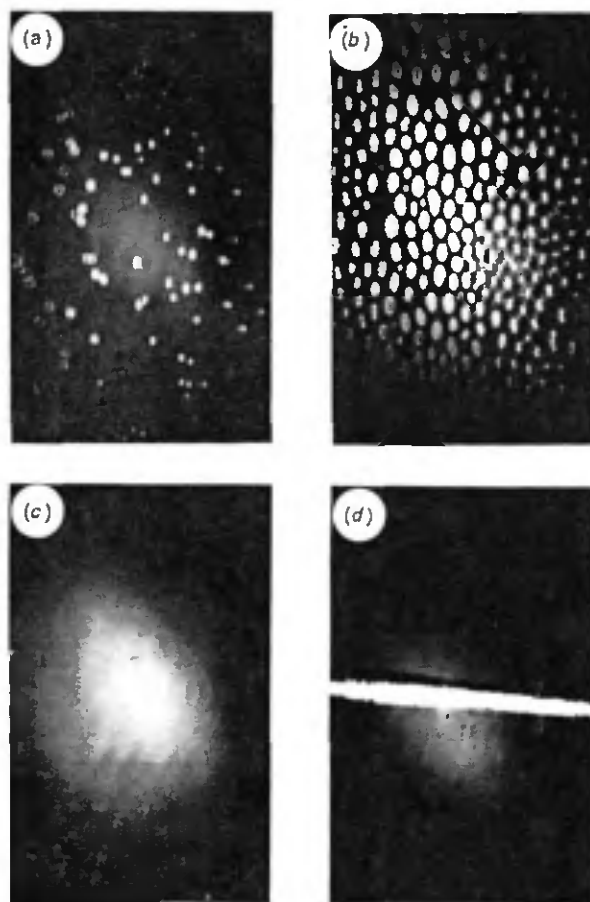


Fig. 6 Brewster-angle microscopy during the compression of cluster 1 [see Fig. 5(b), for example]: circular domains at (a) 1.5 and (b) 0.8 nm^2 ; (c) homogeneous Langmuir film at 0.41 nm^2 (20 mN m^{-1}) (interference fringes are always visible); (d) collapse at 0.30 nm^2 . The images are untreated; owing to the angle of the reflected laser beam^{23b} the domains appear as ellipsoids

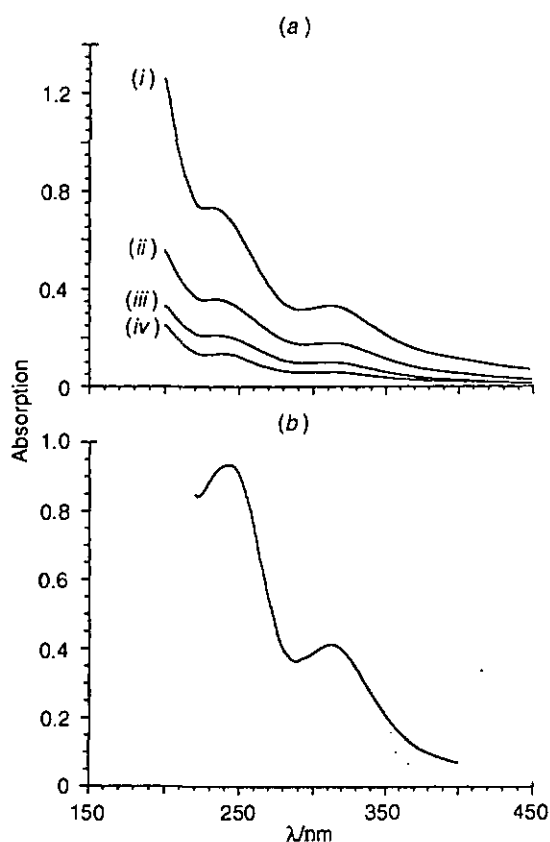


Fig. 7 The UV spectra of cluster 1 transferred onto quartz plates (a) and in CH_2Cl_2 solution ($1.60 \times 10^{-4} \text{ mol dm}^{-3}$) (b). Number of layers in (a): 32 (i), 16 (ii), 8 (iii) and 4 (iv)

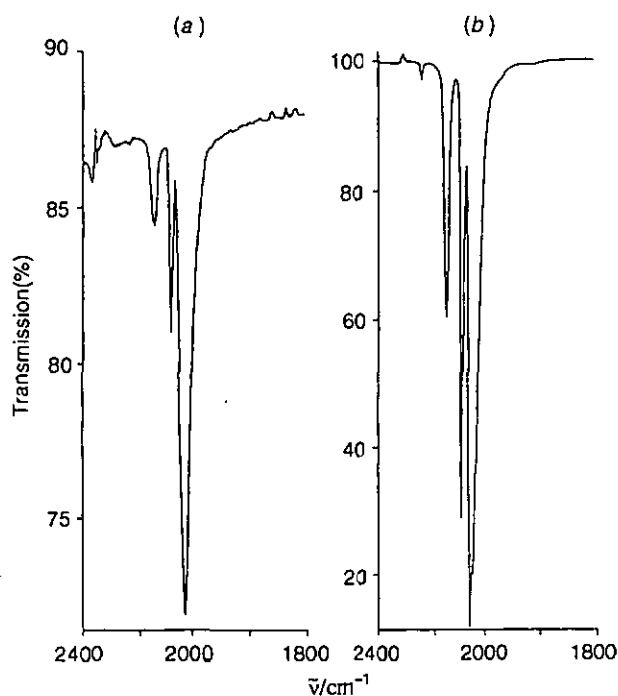


Fig. 8 The IR spectra of cluster 1 transferred onto a silicon wafer (12 transfers, 50 FTIR scans) (a) and in CH_2Cl_2 solution (b)

mN m^{-1} . In each case, identical π vs. A isotherms and BAM images were obtained, demonstrating the reversibility of the overall process.

A detailed interpretation of the above images cannot be given due to the lack of literature data. So far, only fatty acids,²³ phospholipids,^{22a} and more recently a smectic liquid crystalline

polymer²⁴ have been studied by means of BAM during the formation of mono- and/or multi-layers at the air-water interface. In the present study, BAM demonstrated unequivocally the formation of a regular and compact film at 0.46 nm^2 , and confirmed visually the capability of cluster 1 to organize at the air-water interface. In addition, the evolution of the BAM images, with respect to the π and ΔV vs. A isotherms, seems to indicate that the head-to-tail organization of 1 starts from the bright circular domains, the assembling of which leads to the Langmuir films at 0.46 nm^2 .

Y-Type Transferred Films.—Langmuir-Blodgett films were prepared using the conventional dipping method (see Experimental section). The Y-transfer mode (downward and upward deposition) was successfully applied onto different hydrophobized substrates (quartz, glass, mica, silicon wafer) and in each case a transfer ratio (T_R) close to unity was obtained. The transferred films were characterized by means of UV and IR spectroscopy, atomic force microscopy (AFM) and small-angle X-ray scattering (SAXS).

(a) **UV and IR spectroscopy.** Films were deposited onto quartz plates (UV) and silicon wafers (IR), both pretreated with SiMe_2Cl_2 . The UV and IR spectra of the transferred films are identical to those of cluster 1 in solution (Figs. 7 and 8). The UV absorption of the transferred films increased linearly with the number of layers. The chemical stability of the transferred films was evaluated by measuring the variation of the absorption maximum as a function of time. Only a slight decrease of the absorption (about 5%) was detected after 3 weeks, which indicated a good chemical stability of the deposited layers.

(b) **Atomic force microscopy.** Structural properties of films built up from cluster 1 were investigated by means of AFM. This technique²⁵ was successfully applied to image films, obtained from either amphiphilic²⁶ or hydrophobic²⁷ compounds, with molecular resolution.

Two AFM images of five-layer films, transferred onto silicon wafers silanized with SiMe_2Cl_2 , are presented in Fig. 9. At low magnification [Fig. 9(a)] a long-range crystalline ordering of the films is observed. Calculations from the Fourier spectrum gave repeat distances of 0.96 ± 0.05 and 0.95 ± 0.02 nm in the 25 ± 2 and $103 \pm 3^\circ$ directions (measured from the positive x axis). From these values the area of a unit cell is found to be $0.88 \pm 0.06 \text{ nm}^2$. At a higher magnification [Fig. 9(b)] the same values were obtained (within 5%) with the Fourier spectrum showing again the peaks of a periodic structure. Therefore the crystalline structure of the films was confirmed both at low and high magnification.

The cell dimensions calculated by AFM are consistent with the molecular dimensions obtained from the CSC Chem3D PlusTM models. The molecular organization of cluster 1, viewed from the cyano group, is presented in Fig. 10, which images a portion of the surface of the film. Such an arrangement represents a two-dimensional rhombohedral system.

(c) **Small-angle X-ray scattering.** The SAXS experiments were carried out on powders of compounds 1, V and 5 and on films of 1 deposited on glass slides hydrophobized with SiMe_2Cl_2 .

The results for powders of amphiphilic tricobalt cluster 1 showed a well defined diffraction peak at 3.00 ± 0.05 nm. The same value, but with lower intensity, was found, within experimental error, for the cobalt cluster-free cholesterol derivative V. In both cases, wide-angle X-ray scattering (WAXS) results indicated that the peak at 3.0 nm arose from reflections generated by unit cells of monoclinic symmetry. Further studies are planned to estimate the crystal structure of 1 *via* Rietveld refinement of powder data to complete the interpretation of the diffraction peaks. The distance of 3.0 nm corresponds notably to the length of compound V as determined from Corey-Pauling-Koltun (CPK) molecular models. The tricobalt cluster 5, which does not contain the

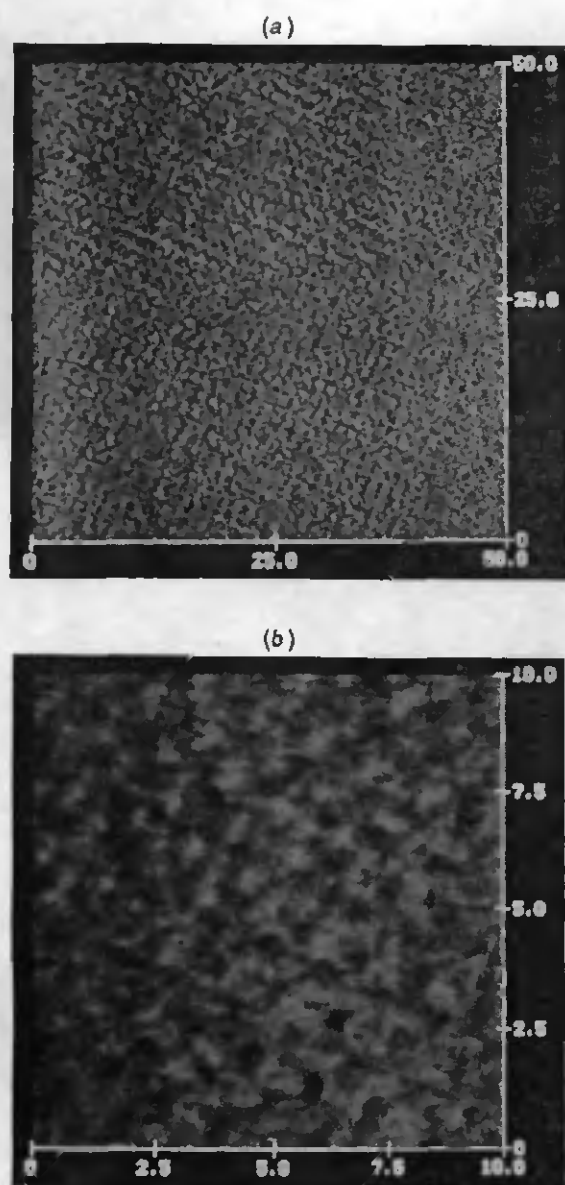


Fig. 9 The AFM images of a five-layer Langmuir-Blodgett film of cluster 1 (low areas are dark, high areas are light) at (a) low (image size 50×50 nm, height 0.5 nm) and (b) high (image size 10×10 nm, height 0.4 nm) magnification. Both scales in nm

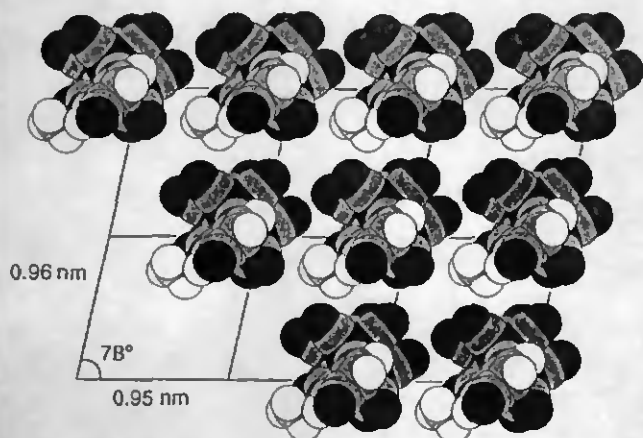


Fig. 10 Proposed molecular organization of cluster 1 of the film surface according to the rhombohedral system provided by the AFM images (viewed from the nitrile group)

cholesterol framework, exhibited no diffraction signal. Furthermore, analysis of the powders by WAXS showed increased disorder and imperfection with increased structural complexity.

The SAXS structure of Langmuir-Blodgett films showed a sharp and intense peak at 6.0 ± 0.1 nm and a much smaller one at 3.1 nm (Fig. 11). The periodicity of 6.0 nm was attributed to the distance between two tricobalt cluster frameworks in the head-to-tail multilayer structure presented in Fig. 12(b). The distance of 3.1 nm, which can be explained by second-order reflection of the periodicity of 6.0 nm, corresponds approximately to the length of the cholesterol derivative [2.9 nm, see Fig. 12(a)], as already noted for the powder analyses. The SAXS

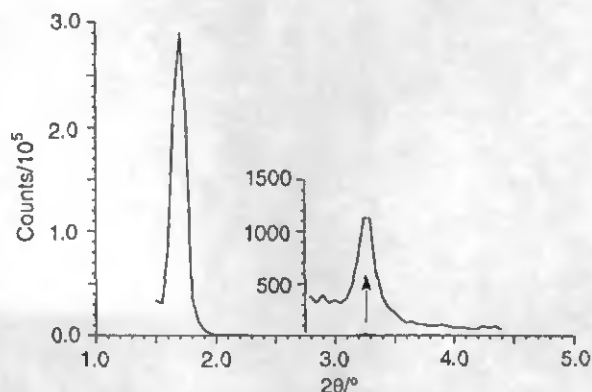


Fig. 11 Small-angle X-ray scattering of a 18-layer film of cluster 1 on a hydrophobic glass plate. High-intensity peak: $2\theta = 1.70^\circ$ ($d = 6.02$ nm). Weak intensity peak (insert): $2\theta = 3.27^\circ$ ($d = 3.13$ nm)

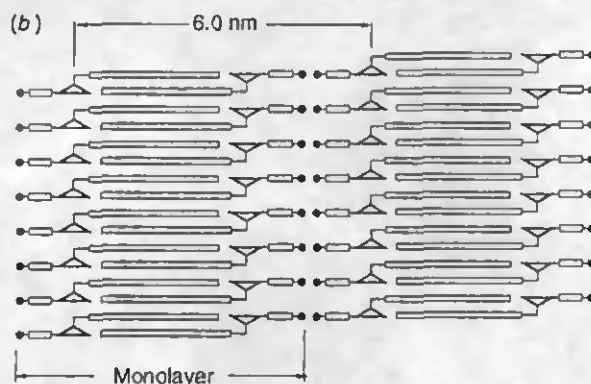
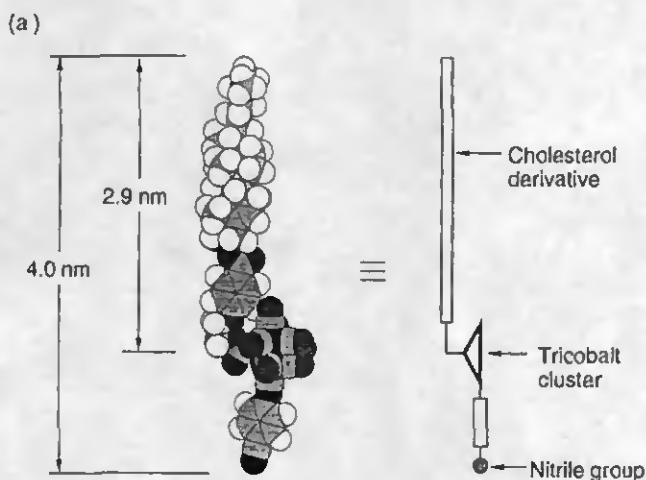


Fig. 12 (a) CSC Chem3D Plus™ structure of compound 1 (drawn with the help of the X-ray data for cluster 4 and those of cholest-5-en- 3β -yl *p*-ethoxybenzoate)²⁸ as well as a symbolic representation. (b) Representation of a bilayer film showing the head-to-tail organization of molecular unit 1 within each monolayer; the distance 6.0 nm was observed from SAXS measurements (see text for further details)

results could only be interpreted from ≥ 12 layers. Indeed, less than 12 layers gave films of low stability, the structure of which was lost within a few hours at room temperature.

The SAXS results conclusively demonstrated that highly ordered films were obtained from cluster I. The head-to-tail structure was observed for all multilayers containing at least 12 layers. The latter were stable for many weeks at room temperature.

Additional Transfer Experiments.—The particular structure of the Langmuir–Blodgett films [see Fig. 12(b)] prompted us to perform other types of transfers. These preliminary experiments, carried out under identical conditions (20 °C, aqueous CdCl_2 – NaHCO_3 subphase), are described below.

(a) A Y-type structure onto a hydrophobic substrate (glass treated with SiMe_2Cl_2) could also be obtained ($T_R \approx 1$) starting from an upward deposition.

(b) Films of cluster I were prepared onto a hydrophobic glass (SiMe_2Cl_2) using the X-type transfer mode (downward deposition). A 12-layer film was obtained with excellent transfer ratios ($T_R \approx 1$). Obviously, if head-to-tail layers are transferred, the X- and Y-deposited films should have the same structure. However, it is important to mention that the experimental procedure applied in these two different processes may influence the film organization (during the X-transfer mode the coated plates are allowed to dry before subsequent deposition).

(c) One monolayer could be deposited onto a hydrophilic glass plate by lifting up the substrate from the subphase ($T_R \approx 1$). However, the subsequent dipping caused a quantitative return of the transferred monolayer into the floating film.

These results strengthen the evidence for the film structure illustrated in Fig. 12(b). First, experiments (a) and (b) are in agreement with the head-to-tail molecular organization within the monolayer. Secondly, related to (a) and (b), experiment (c) points out the amphiphilic nature of the monolayer surface. Nevertheless, the molecular organization of the first monolayer as well as its interaction with the solid substrates are still unknown. X-Ray measurements have shown that films constituted of few monolayers do not exhibit a well defined structure. The head-to-tail organization presented in Fig. 12(b) takes shape during the increasing number of transfers. Ongoing research to elucidate the remaining questions is in progress.

Conclusion

Stable layers at the air–water interface and ordered Langmuir–Blodgett films were successfully obtained from an amphiphilic tricobalt cluster derivative. The supramolecular organization of the tricobalt cluster-containing molecular unit was carefully investigated by means of surface pressure and surface potential *vs.* area isotherms, Brewster-angle and atomic force microscopy and small-angle X-ray scattering. The postulated head-to-tail structure is consistent with all the data obtained from the above experimental techniques.

The major interest of the work described herein consists of the possibility to design and synthesise organometallic solids and surfaces from the organization of polynuclear transition metal-containing building blocks. Such organometallic assemblies might exhibit new magnetic, electrochemical and optical properties. They could also be used in heterogeneous catalytic reactions and to investigate new substrate–adsorbate interactions in solid–gas adsorption processes (chemical adsorption of CO).

Experimental

General.—Clusters **3**¹⁷ and **5**¹⁶ were prepared according to literature procedures. All other chemicals were used as received without further purification. Thermal properties of cholesterol derivatives III–V will be reported elsewhere. Column chroma-

tography: silica gel 60 (0.063–0.200 mm, Merck). TLC: silica gel plates 60 F₂₅₄ (Merck). Melting points: Büchi 510 instrument (uncorrected). α : Perkin Elmer 241 polarimeter. UV spectra: Uvikon 930 spectrophotometer. IR spectra: Perkin Elmer 1720 FTIR spectrometer. ¹H and ¹³C NMR spectra: Bruker AMX 400 spectrometer at 400.13 MHz (¹H) and 100.62 MHz (¹³C); SiMe_4 as internal reference; L refers to co-ordinated compound VIII. Mass spectra (chemical ionization, NH_3): Nermag R 30.10 spectrometer. Program CSC Chem 3D Plus™. Cambridge Scientific Computing, USA.

Crystal Structure Determination of Cluster 4.— $\text{C}_{19}\text{H}_7\text{Co}_3\text{N}_2\text{O}_{10}$, $M_r = 600.1$, triclinic, space group $P\bar{1}$, $a = 9.165(2)$, $b = 9.455(2)$, $c = 14.107(2)$ Å, $\alpha = 70.69(1)$, $\beta = 73.78(1)$, $\gamma = 84.83(1)^\circ$, $U = 1107.7$ Å³, $Z = 2$, $D_c = 1.799$ g cm⁻³, $\lambda = 0.71073$ Å, $\mu = 22.7$ cm⁻¹, $F(000) = 594$.

3899 Unique reflections, 3245 observed [$I > 3\sigma(I)$], $R = 0.046$, $R' = 0.067$, $k = 0.002$, $S = 1.40$. Maximum shift/error ratio 0.005:1, residual density (e Å⁻³) maximum 0.74, minimum -1.19.

Intensity data were collected at room temperature on a Stoe AED2 four-circle diffractometer using Mo-K α graphite-monochromated radiation. The structure was solved by Patterson and Fourier difference syntheses using the NRCVAX²⁹ system, which was used for all further calculations. Neutral complex-atom scattering factors used in NRCVAX²⁹ were from ref. 30. The H atoms were included in calculated positions [$U_{iso} = U_{eq}(C) + 0.01$ Å²] and renewed after every third cycle of refinement. The non-hydrogen atoms were refined anisotropically using weighted full-matrix least squares, where $w = 1/[\sigma^2(F_o) + k(F_o^2)]$. The numbering scheme used is apparent from Fig. 1, drawn using the program PLUTO.¹⁹

Additional material available from the Cambridge Crystallographic Data Centre comprises H-atom coordinates, thermal parameters and remaining bond lengths and angles.

Langmuir–Blodgett Technique.—Instrumentation: KSV Instrument Inc., Finland. Experiments were carried out on a Langmuir trough LB 5000 (Teflon™ coating, dimensions: 50 × 15 cm) equipped with two hydrophobic barriers, a Wilhelmy balance as a surface-pressure sensor, a film-deposition system and a 5000 SP surface-potential meter (vibrating capacitor method). The whole system, fully computer controlled, lay on an active vibration-isolating table and was protected from dust by a home-made plastic cover.

Measurements of π *vs.* A isotherms were carried out by the ordinary method. Typical conditions were as follows: subphase, water (18 M Ω cm⁻¹; Elgastat UHQ-II, Kleiner, Switzerland) or aqueous CdCl_2 (1.5×10^{-4} mol dm⁻³) and NaHCO_3 (5×10^{-3} mol dm⁻³) solution; temperature, subphase thermostatted within ± 0.2 °C, atmosphere at 20–22 °C; concentration and volume of chloroform solution of the amphiphile, 1 g dm⁻³, 60–100 mm³; compression speed, 30 cm² min⁻¹. Vertical transfers were performed under the following conditions: subphase, aqueous CdCl_2 (1.5×10^{-4} mol dm⁻³) and NaHCO_3 (5×10^{-5} mol dm⁻³) solution maintained at 20.0 ± 0.2 °C; deposition surface pressure, 20 mN m⁻²; deposition speed, 0.5 cm min⁻¹.

Brewster-angle Microscopy.—Instrumentation: BAM-1 Brewster-angle microscope (Nanofilm Technologie, Germany). A detailed description of the instrument was recently published.^{23b}

Atomic Force Microscopy.—Instrumentation: Nanoscope II AFM (Digital Instruments, USA). Procedure: five layers of cluster I were transferred onto a previously silanized (SiMe_2Cl_2) ultraflat silicon wafer (Si{100}; Faseltec, Switzerland). The next day the wafer was glued to a steel plate and set up into the contact atomic force microscope. Imaging was done

in water with a scanner of 14 μm range and with a short Olympus tip. Only spike removal and low-pass filtering were applied to the images. The AFM analysis of the silanized silicon wafers gave no structure. Further details can be found elsewhere.²⁷

Small-angle X-Ray Scattering.—All measurements were performed using the Stoe-Stadi P powder diffractometer system. Focused $\text{Co-K}\alpha_1$ radiation with $\lambda = 0.178\ 898\ \text{nm}$ was obtained by a curved $\text{Ge}(111)$ monochromator. The powdered samples were investigated in transmission mode with a rotating-sample holder to minimize preferred orientation of the plate-like crystallites. Data were collected with a linear position-sensitive detector (Stoe) in the $1.5\text{--}7^\circ$ (2θ) range (and up to 120° for wide-angle X-ray scattering).

The SAXS results for Langmuir–Blodgett films were recorded using a sample holder for reflection mode in the 2θ range $0.5\text{--}5^\circ$ in combination with a scintillation counter at 130 mm from the sample and with a scan width of 0.02° . The instrument was carefully adjusted by measuring multilayers of cadmium eicosanoate from which good agreement with published data could be achieved.

Syntheses.—**Methyl 4-methoxymethoxybenzoate II.** Dimethoxymethane (75.2 g, 1.00 mol) and P_2O_5 (28.0 g, 0.20 mol) were added to methyl 4-hydroxybenzoate I (15.0 g, 0.10 mol) in dry CH_2Cl_2 (400 cm^3). The mixture was stirred at room temperature overnight, filtered and evaporated. The residue was dissolved in Et_2O (300 cm^3) and the ether phase was washed with saturated aqueous Na_2CO_3 and saturated aqueous NaCl , dried over MgSO_4 and evaporated. The residue was purified by column chromatography (MeCO_2Et –hexane, 1:5) to afford compound II (17.50 g, 89%) as a colourless liquid, R_f (MeCO_2Et –hexane, 1:1) 0.52 (Found: C, 61.1; H, 6.3. $\text{C}_{10}\text{H}_{12}\text{O}_4$ requires C, 61.2; H, 6.2%); $\nu_{\text{max}}(\text{CH}_2\text{Cl}_2)$ 2985–2831, 1733 and $1607\ \text{cm}^{-1}$; $\delta_{\text{H}}(\text{CDCl}_3)$ 3.48 (3 H, s, OCH_3), 3.88 (3 H, s, CO_2CH_3), 5.22 (2 H, s, OCH_2O), 7.05 (2 H, d, aromatic H) and 7.98 (2 H, d, aromatic H).

Cholest-5-en-3 β -yl 4-methoxymethoxybenzoate III. Butyllithium (1.6 mol dm^{-3} in hexane, 100 cm^3) was added dropwise, under N_2 , to a solution of cholesterol (59.1 g, 0.15 mol) in dry thf (150 cm^3). The mixture was stirred at 0°C for 15 min. A solution of methyl ester II (15.0 g, 0.08 mol) in dry thf (100 cm^3) was added dropwise and the mixture was stirred at room temperature overnight and evaporated. The residue was dissolved in Et_2O (700 cm^3) and the organic phase was washed with saturated aqueous NaCl ($3 \times 200\ \text{cm}^3$), dried (MgSO_4) and evaporated. The residue was purified by column chromatography (MeCO_2Et –hexane, 1:6) to give compound III (35.10 g, 84%) as a white solid, R_f (MeCO_2Et –hexane, 1:1) 0.68, $\alpha -4.7^\circ$ (589.3 nm, 22°C , $c\ 4.27 \times 10^{-2}\ \text{g cm}^{-3}$, CHCl_3) (Found: C, 78.5; H, 9.9. $\text{C}_{36}\text{H}_{54}\text{O}_4$ requires C, 78.5; H, 10.0%); $\nu_{\text{max}}(\text{KBr})$ 2945–2866, 1717, 1607 and $1275\ \text{cm}^{-1}$; $\delta_{\text{H}}(\text{CDCl}_3)$ 0.69–2.44 (43 H, cholesteryl), 3.48 (3 H, s, OCH_3), 4.82 (1 H, m, CHO, cholesteryl), 5.23 (2 H, s, OCH_2O), 5.42 (1 H, m, C=CH, cholesteryl), 7.05 (2 H, d, aromatic H) and 7.98 (2 H, d, aromatic H); chemical ionization (CI) mass spectrum m/z 552 ($[M^+ + 2]$, 1), 369 (86), 368 (100), 353 (18), 283 (2), 247 (10), 213 (14), 165 (37), 147 (60) and 121 (37%).

Cholest-5-en-3 β -yl 4-hydroxybenzoate IV. A mixture of compound III (3.70 g, 6.64 mmol), MeCO_2H (100 cm^3), 37% aqueous HCl (10 cm^3) and thf (70 cm^3) was stirred at room temperature for 8 h. The desired product, which precipitated during the reaction, was filtered off and washed with thf ($2 \times 20\ \text{cm}^3$) to afford compound IV (2.88 g, 85%) which was used in the following step without further treatment. Purification by column chromatography (Et_2O – CH_2Cl_2 –hexane, 1:15:8) followed by two crystallizations from toluene gave analytically pure IV, R_f (Et_2O –hexane, 2:1) 0.43, $\alpha -6.3^\circ$ (589.3 nm, 23°C , $c\ 0.40 \times 10^{-2}\ \text{g cm}^{-3}$, CHCl_3) (Found: C, 80.6; H, 9.9. $\text{C}_{34}\text{H}_{50}\text{O}_3$ requires C, 80.4; H, 9.9%); $\nu_{\text{max}}(\text{KBr})$ 3383, 2950–

2866, 1711, 1608, 1274 and $1242\ \text{cm}^{-1}$; $\delta_{\text{H}}(\text{CDCl}_3)$ 0.68–2.45 (43 H, cholesteryl), 4.81 (1 H, m, CHO, cholesteryl), 5.30 (1 H, s, OH), 5.42 (1 H, m, C=CH, cholesteryl), 6.86 (2 H, d, aromatic H) and 7.97 (2 H, d, aromatic H).

Cholest-5-en-3 β -yl 4-(2-hydroxyethoxy)benzoate V. A mixture of phenol IV (2.00 g, 3.75 mmol), 2-bromoethanol (1.41 g, 11.3 mmol) and K_2CO_3 (1.56 g) in dry dimethylformamide (dmf) (80 cm^3) and dry thf (50 cm^3) was stirred at 120°C for 14 h. Additional 2-bromoethanol (1.41 g, 11.3 mmol) and K_2CO_3 (1.56 g) were added, and the mixture stirred at 120°C for 3 h. After cooling to room temperature, K_2CO_3 was filtered off and the solution evaporated. The residue was purified by column chromatography (Et_2O – CH_2Cl_2 –hexane, 2:1:1) to give compound V (1.71 g, 83%) as a white solid, R_f (Et_2O) 0.34, $\alpha -3.4^\circ$ (589.3 nm, 22°C , $c\ 4.41 \times 10^{-2}\ \text{g cm}^{-3}$, CHCl_3) (Found: C, 78.5; H, 9.9. $\text{C}_{36}\text{H}_{54}\text{O}_4$ requires C, 78.5; H, 9.9%); $\nu_{\text{max}}(\text{KBr})$ 3543, 2954–2867, 1712, 1609, 1384 and $1276\ \text{cm}^{-1}$; $\delta_{\text{H}}(\text{CDCl}_3)$ 0.69–2.45 (43 H, cholesteryl), 3.98 (2 H, m, $\text{HOCH}_2\text{CH}_2\text{O}$), 4.14 (2 H, t, $\text{HOCH}_2\text{CH}_2\text{O}$), 4.83 (1 H, m, CHO, cholesteryl), 5.41 (1 H, m, C=CH, cholesteryl), 6.93 (2 H, d, aromatic H) and 8.00 (2 H, d, aromatic H); CI mass spectrum m/z 386 (4), 369 (35), 368 (48), 353 (7), 247 (9), 213 (8), 165 (47), 147 (39) and 121 (100%).

Cluster 2. Freshly sublimed AlCl_3 (1.26 g, 9.45 mmol) was added, under N_2 , to a solution of $[\text{Co}_3(\text{CCl})(\text{CO})_9]$ 3 (1.50 g, 3.15 mmol) in dry CH_2Cl_2 . After stirring at room temperature for 30 min the mixture, initially purple, turned brown and a precipitate formed; TLC (hexane) indicated complete consumption of 3. Compound V (1.56 g, 2.84 mmol) was added. The purple solution was stirred for 10 min and evaporated. The residue was dissolved in Et_2O (200 cm^3) and the ether phase was washed twice with 5% aqueous HCl , dried (MgSO_4) and evaporated. The solid residue was purified by column chromatography (CH_2Cl_2 –hexane, 1:1) and crystallized from hexane at -20°C to give the tricobalt cluster 2 (1.27 g, 44%) as a purple solid, R_f (Et_2O –hexane, 1:1) 0.54, m.p. 156°C (decomp.) (Found: C, 55.8; H, 5.2. $\text{C}_{47}\text{H}_{53}\text{Co}_3\text{O}_{14}$ requires C, 55.4; H, 5.2%); $\nu_{\text{max}}(\text{CH}_2\text{Cl}_2)$ 2111w, 2065vs and $2047\ \text{cm}^{-1}$; $\delta_{\text{H}}(\text{CDCl}_3)$ 0.69–2.44 (43 H, cholesteryl), 4.29 (2 H, t, $\text{CO}_2\text{CH}_2\text{CH}_2\text{O}$), 4.67 (2 H, t, $\text{CO}_2\text{CH}_2\text{CH}_2\text{O}$), 4.81 (1 H, m, CHO, cholesteryl), 5.41 (1 H, m, C=CH, cholesteryl), 6.87 (2 H, d, aromatic H) and 7.98 (2 H, d, aromatic H); $\delta_{\text{C}}(\text{CDCl}_3)$ 64.27 (OCH_2), 66.70 (OCH_2), 114.57 (2 C, aromatic CH), 124.42 (aromatic CCO_2), 132.24 (2 C, aromatic CH), 162.73 (aromatic CO), 166.35 (CO_2 of cholesteryl), 179.40 (CO_2CH_2), 199.17 (9 C, CO), 252.06 (CCO_2) and 12.56, 19.42, 20.06, 21.76, 23.24, 23.49, 24.53, 24.99, 28.64, 28.70, 28.92, 32.63, 32.65, 36.49, 36.89, 37.36, 37.76, 38.98, 40.22, 40.46, 43.03, 50.77, 56.86, 57.41, 74.97, 123.37 and 140.45 (27 C, cholesteryl).

4-Formylaminobenzonitrile X. A mixture of 4-aminobenzonitrile IX (5.00 g, 0.042 mol) and HCO_2H (30 cm^3) was heated at reflux for 3 h. After cooling to room temperature the solution was evaporated. Recrystallization of the residue from water (100 cm^3) afforded 4-formylaminobenzonitrile X (4.63 g, 75%) as a white solid m.p. 189°C (lit.,³¹ $188\text{--}189^\circ\text{C}$); $\nu_{\text{max}}(\text{KBr})$ 3320–2800, 2223, 1703, 1608, 1523 and $1494\ \text{cm}^{-1}$.

4-Isocyanobenzonitrile VIII. The compound POCl_3 (4.18 g, 0.027 mol) was added dropwise, under N_2 , to a mixture of 4-formylaminobenzonitrile X (3.65 g, 0.025 mol), dry diisopropylamine (6.82 g, 0.067 mol) and dry CH_2Cl_2 (100 cm^3), cooled to 0°C . The mixture was stirred at 0°C for 1 h and at room temperature for 14 h. A solution of Na_2CO_3 (4.9 g, 0.046 mol) in water (30 cm^3) was added dropwise, and the mixture stirred for 1 h. Dichloromethane (50 cm^3) was added and the organic phase separated, washed with saturated aqueous NaHCO_3 and saturated aqueous NaCl , dried (MgSO_4) and evaporated. The residue was purified by column chromatography (CH_2Cl_2) to give 4-isocyanobenzonitrile VIII (2.60 g, 80%) as a white, bad-smelling solid, R_f (CH_2Cl_2) 0.60, m.p. 129°C (decomp.) (lit.,³² 130°C) (Found: C, 74.5; H, 3.2; N, 21.75. Calc. for $\text{C}_8\text{H}_4\text{N}_2$: C, 75.0; H, 3.15; N, 21.85%); $\nu_{\text{max}}(\text{KBr})$ 3096, 3048, 2232, 2133,

1605 and 1500 cm^{-1} ; $\delta_{\text{H}}(\text{CDCl}_3)$ 7.50 (2 H, d, aromatic H) and 7.73 (2 H, d, aromatic H); $\delta_{\text{C}}(\text{CDCl}_3)$ 113.46 (C^1), 117.12 (PhCN), 127.27 (3 C, C^{3-5}), 133.50 (2 C, $\text{C}^{2,6}$) and 169.20 (PhNC).

Cluster 1. Isocyanide VIII (63 mg, 0.49 mmol) was added, under N_2 , to a solution of cluster 2 (200 mg, 0.20 mmol) in dry thf (15 cm^3). The mixture was stirred at room temperature for 5 h, filtered through a short silica gel column and eluted with CH_2Cl_2 . The mixture was evaporated and the residue submitted to column chromatography (hexane- CH_2Cl_2 - Et_2O , 6:2:1) to separate unreacted 2 (35%), which eluted first, followed by cluster 1 (27%) and then by polysubstituted clusters. Recrystallization at -20°C afforded pure I (55 mg, 25%) as a purple solid, R_f (Et_2O -hexane, 1:1) 0.40, m.p. $< 50^\circ\text{C}$ (decomp.) (Found: C, 57.75; H, 5.05; N, 2.55. $\text{C}_{54}\text{H}_{57}\text{Co}_3\text{N}_2\text{O}_{13}$ requires C, 58.0; H, 5.1; N, 2.5%; $\lambda_{\text{max}}/\text{nm}$ ($\epsilon/\text{dm}^3\text{mol}^{-1}\text{cm}^{-1}$) (CH_2Cl_2) 243 (58 300), 312 (25 800) and 540 (1900); $\nu_{\text{max}}(\text{CH}_2\text{Cl}_2)$ 2235m, 2142s, 2087vs and 2054 (sh); (KBr) 3000-2800, 2233, 2152, 2092, 2045, 1708, 1673 and 1604 cm^{-1} ; $\delta_{\text{H}}(\text{CDCl}_3)$ 0.66-2.46 (43 H, cholesteryl), 4.27 (2 H, t, $\text{CO}_2\text{CH}_2\text{CH}_2\text{O}$), 4.65 (2 H, t, $\text{CO}_2\text{CH}_2\text{CH}_2\text{O}$), 4.82 (1 H, m, CHO, cholesteryl), 5.42 (1 H, m, $\text{C}=\text{CH}$, cholesteryl), 6.83 [2 H, d, aromatic H (benzoyl)], 7.43 [2 H, d, aromatic H (L)], 7.66 [2 H, d, aromatic H (L)] and 7.94 [2 H, d, aromatic H (benzoyl)]; $\delta_{\text{C}}(\text{CDCl}_3)$ 63.96 (OCH_2), 66.92 (OCH_2), 113.97 (aromatic CCN), 114.58 [2 C, aromatic CH (benzyl)], 117.81 (PhCN), 124.40 (aromatic CCO_2), 127.63 [2 C, aromatic CH (L)], 131.76 (aromatic CNC), 132.21 [2 C, aromatic CH (benzoyl)], 134.26 [2 C, aromatic CH (L)], 162.77 (aromatic CO), 163.79 (PhNC), 166.29 (CO_2 of cholesteryl), 180.56 (CO_2CH_2), 200.74 (8 C, CO), 252.03 (CCO_3) and 12.55, 19.42, 20.06, 21.75, 23.24, 23.49, 24.53, 24.99, 28.64, 28.70, 28.92, 32.60, 32.62, 36.49, 36.89, 37.36, 37.75, 38.98, 40.21, 40.46, 43.03, 50.77, 56.86, 57.41, 75.04, 123.40 and 140.43 (27 C, cholesteryl).

[$\text{Co}_3(\text{CO})_8(\text{CCO}_2\text{Me})(\text{CNC}_6\text{H}_4\text{CN})$] 4. This compound was obtained by reaction of tricobalt cluster 5 (500 mg, 1.00 mmol) with isocyanide VIII (315 mg, 2.50 mmol) as described above for compound 1. Separation by column chromatography (hexane- CH_2Cl_2 - Et_2O , 4:1:1) gave unreacted 5 (70%), which eluted first, followed by cluster 4 (23%) and then by polysubstituted clusters. Recrystallization from hexane- CH_2Cl_2 at -20°C afforded pure 4 (134 mg, 22%) as a purple solid, R_f (Et_2O -hexane, 1:1) 0.36, m.p. $< 50^\circ\text{C}$ (decomp.) (Found: C, 37.9; H, 1.2; N, 4.6. $\text{C}_{19}\text{H}_7\text{Co}_3\text{N}_2\text{O}_{10}$ requires C, 38.0; H, 1.2; N, 4.7%; $\nu_{\text{max}}(\text{CH}_2\text{Cl}_2)$ 2236m, 2144s, 2088vs and 2055 (sh) cm^{-1} ; $\delta_{\text{H}}(\text{CDCl}_3)$ 3.87 (3 H, s, CH_3), 7.49 (2 H, d, aromatic H) and 7.72 (2 H, d, aromatic H); $\delta_{\text{C}}(\text{CDCl}_3)$ 53.26 (CH_3), 113.92 (aromatic CCN), 117.90 (PhCN), 127.69 (2 C, aromatic CH), 131.87 (aromatic CNC), 134.30 (2 C, aromatic CH), 164.18 (PhNC), 181.24 (CO_2), 200.85 (8 C, CO) and 253.47 (CCO_3).

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